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ChemQuest 35 Intro, to  
chemical Reoctioil\$"

Date: Hour: Formula

Review 1. Write the  
formulas for the

following compounds:

2. Write the names for  
the folowing ...

### **Intro, to chemical Reoctioil\$ - LTHS**

### **Answers**

14. Based on your  
answer to question 13,  
verify that this

equation is true: This  
equation is called the

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“Combined Gas Law”. Notice that it contains all of the equations combined into one! Again, the pressure times the volume divided by the Kelvin temperature equals 1.18 for every trial in the table. 15.

### **Information: Gas Pressure**

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Booth, Powder Coating  
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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW

Practice Keys -

ChemActivity 51: Cell Voltage - ChemActivity

50: Electrochemical

Cell UNIT 11 - HW

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ChemQuest 55: Free Energy - ChemQuest

54: 2nd Law of

Thermodynamics

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**HW Keys - Roosevelt  
High School AP  
Chemistry 2017-18**

ChemQuest 20 Name:

\_\_\_\_\_ Date: \_\_\_\_\_ Hour:

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Polyatomic Ions The word, "polyatomic" means "many atoms".

A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion,  $\text{SO}_4^{2-}$ .

Sulfate is composed

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#### **Information :**

#### **Polyatomic Ions**

concept of significant figures has the answer.

When multiplying and dividing numbers, you need to round your answers to the correct number of significant figures. To round correctly, follow these simple steps: 1) Count the number of significant figures in each number. 2) Round your answer to the

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least number of  
significant figures.

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A. Hough High School.  
I. Single Replacement  
Reactions 1. Write the  
reaction for each of the  
following reactions: a)  
Sodium sulfate reacts

**Chemquest 35 - I**  
*Page 15/24*

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the ...**

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*Page 19/24*

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**Stanford University**

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Date: Hour Information:

Double Replacement

Reactions In the

previous ChemQuest,

you learned about

single replacement

reactions in which a

single atom replaces

an ion from another

reactant. Study what

happens in the

following reactions.

They are called double

replacement reactions.

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**Fig-t - LTHS Answers**

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*Page 22/24*

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