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NH 4 OH (aq) -> NH 4 + (aq) + OH - (aq) Ca (OH) 2 (aq) -> Ca 2 + (aq) + 2OH - (aq) Mg (OH) 2 (aq) -> Mg 2 + (aq) + 2OH - (aq) 3. An acid is a proton donor. A base is a proton acceptor. this is called Bronsted-Lowry definition of acids and bases. From this definition, an acid donates H⁺ + .

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Short Notes: Form 4 Chemistry Chemical Formulae and Equation Calculation Summary Summary + molar volume × molar volume × molar mass Avogadro Constant Number of particles The metal oxide of group 1 elements can dissolve in water to form ...

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