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Chapter-14-Chemistry-Answers 2/3 PDF Drive - Search and download PDF files for free. CHAPTER 14 REVIEW Acids and Bases CHAPTER 14 REVIEW Acids and Bases SHORT ANSWER Answer the following questions in the space provided 1 a Write the two equations that show the two-stage ionization of sulfurous acid in water stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$ b Which stage of ionization usually produces more ions?

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1 . One example for NH_3 as a conjugate acid: $\text{NH}_2^- + \text{H}^+ \rightarrow \text{NH}_3$; $\text{NH}_2^- + \text{H}^+ \rightarrow \text{NH}_3$; as a conjugate base: $\text{NH}_4^+ + (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{NH}_3 (\text{aq})$. (a) $\text{NaHSO}_3 < \text{NaHSO}_4$; in polyoxy acids, the more electronegative central element—S, in this case—forms the stronger acid. The larger number of oxygen atoms on the central atom (giving it a higher oxidation state) also creates ...

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Download File PDF Chapter 14 Chemistry Answers Section Assessment 14.1 page 479 1. Explain Use the properties of seawater to describe the characteristics of mixtures. Answers will vary but might include that Mixtures and Solutions Answer Questions, Exercises, and Problems 14–15. Check your answers with those at the end of the chapter.

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CHAPTER 14 REVIEW Acids and Bases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Name the following compounds as acids: sulfuric acid a. H_2SO_4 sulfurous acid b. H_2SO_3 hydrosulfuric acid c. H_2S perchloric acid d. HClO_4 hydrocyanic acid e. hydrogen cyanide 2. H_2S Which (if any) of the acids mentioned in item 1 are binary acids? 3.

14 Acids and Bases - David Brearley High School

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Chapter 14 Chemical Kinetics

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Chemistry 102 ANSWER KEY REVIEW QUESTIONS Chapter 14 1. A mixture of 0.10 mol of NO, 0.050 mol of H₂ and 0.10 mol of H₂O is placed in a 1.0L flask and allowed to reach equilibrium as shown below: $\text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g}) \rightleftharpoons 2 \text{NO}(\text{g}) + 2 \text{H}_2(\text{g})$ At equilibrium $[\text{NO}] = 0.062 \text{ M}$. Calculate the equilibrium constant, K_c, for this reaction.

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