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17 Terms.

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interactions invol....

Identify the type of
interactions invol....

Identify the type of
interactions invol....

The separation of
solute particles (ΔH_1

Solvent-solvent
Interactions...

-interactions between
the water m....

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Chapter 12 Molecular
Structure □ Review
Skills 12.1 A New Look
at Molecules and the
Formation of Covalent
Bonds □ The Strengths
and Weaknesses of
Models □ The Valence-
Bond Model 12.2
Drawing Lewis
Structures □ General
Procedure □ More Than

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Chemistry
One Possible Structure
12.3 Resonance

Answer Key

Chapter 12 Molecular Structure - An Introduction to Chemistry

25 terms. jakeydv.
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chemistry. law of
conservation of mass.
formula mass. law of
definite proportions.
limiting reactant. no
detectable change in
mass occurs during a
chemical reaction. the

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sum of the atomic masses given in the formula of a substan....

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Review Skills 12.1

Change from Gas to
Liquid and from Liquid
to Gas - An

Introduction to
Dynamic Equilibrium
The Process of
Condensation The
Process of Evaporation
How Evaporation
Causes Cooling Rate of
Evaporation Dynamic
Equilibrium Between
Liquid and Vapor
Equilibrium Vapor
Pressure.

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Chapter 12 Liquids: Condensation, Evaporation, and Dynamic ...

Section 12.1 Gases 1.
motion 2. a. small b.
forces c. random d.
elastic; kinetic 3. KE
 $\frac{1}{2}mv^2$ 4.
Temperature 5. true 6.
true 7. false 8. true 9.
true 10. false 11. true
12. false 13. a 14. a 15.
d 16. d 17. b 18. b 19.
b 20. barometer 21.
Evangelista Torricelli
22. gravity and

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atmospheric pressure

23. The atmospheric
pressure has

increased. Section 12

.2 Forces of Attraction

1. intermolecular

forces 2.

intermolecular forces

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States of Matter 30
Terms. Samplesa.

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questions 8 Terms.
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Terms.

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the study of all
substances containing.
Inorganic chemistry.
The branch of
chemistry that deals

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- Chapter 12 - States of
Matter Section 12.1

Gases 1. motion 2. a.

small b. forces c.

random d. elastic;

kinetic 3.

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Alcohols Alcohols are
organic compounds
that have a hydroxyl
(-OH)

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calculation of quantities in chemical reactions. The quantities to consider in a chemical reaction? On the left side of a chemical equation, you have the? On the right side of a chemical equation, you have the? In every chemical reaction, these two things are conserved. In a majority of chemical reactions, these three things are NOT conserved.

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